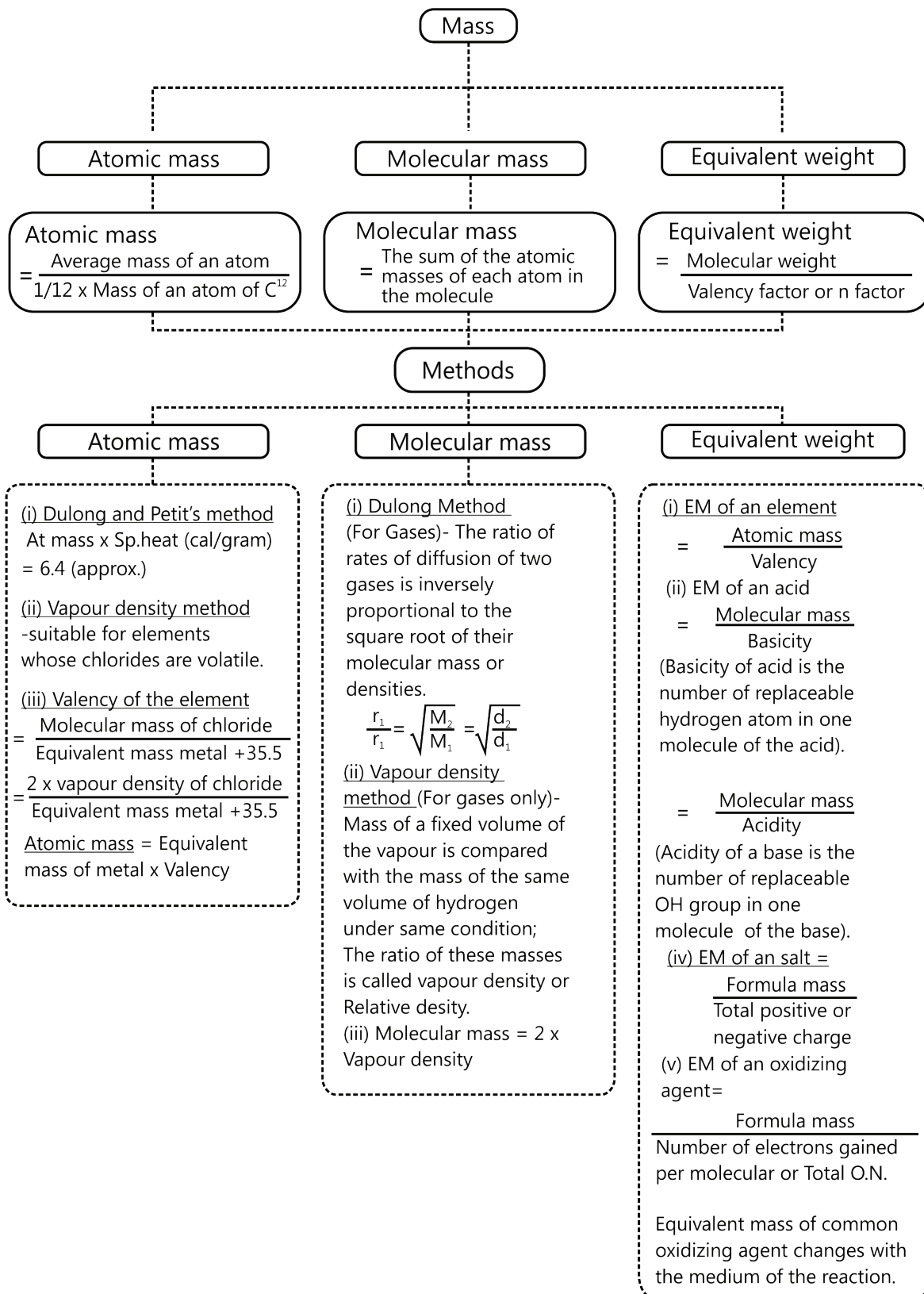
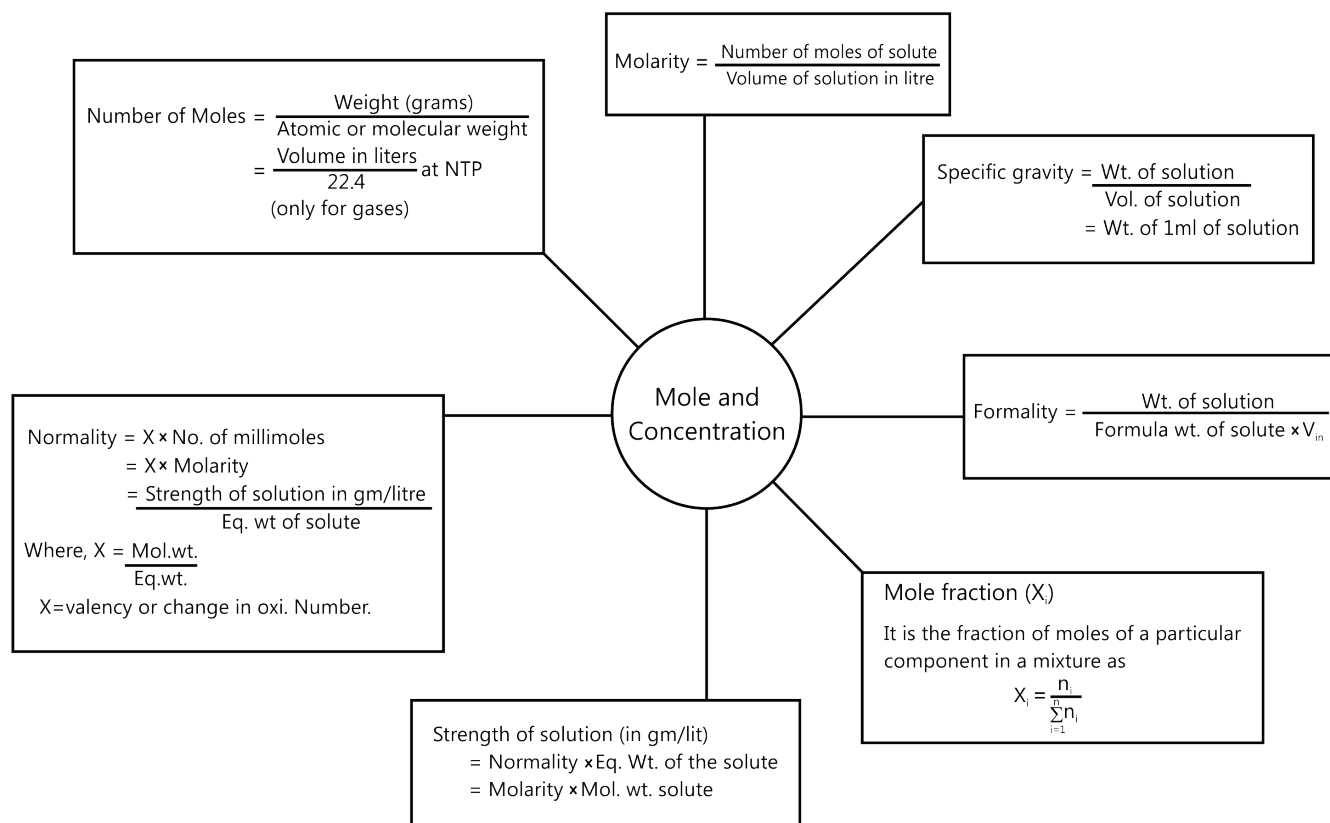


FORMULAE SHEET





RULES IN BRIEF

The following are the definitions of 'mole' represented in the form of equations:

(a) Number of moles of molecules = $\frac{\text{Weight in g}}{\text{Molecular weight}}$

(b) Number of moles of atoms = $\frac{\text{Weight in g}}{\text{Atomic weight}}$

(c) Number of moles of gases = $\frac{\text{Volume at NTP}}{\text{Standard molar volume}}$

(Standard molar volume is the volume occupied by 1 mole of any gas at NTP, which is equal to 22.4 litres.)

(d) Number of moles of atoms / molecules / ions / electrons = $\frac{\text{No. of atoms / molecules / ions / electrons}}{\text{Avogadro constant}}$

(e) Number of moles of solute = Molarity \times Volume of solution in litres

Or No. of millimoles = Molarity \times Volume in mL.

$$\frac{\text{Millimoles}}{1000} = \text{moles}$$

(f) For a compound $M_x N_y$, x moles of N = y moles of M

